Synthesis of π -Extended [2*n* + 1]Helicenes: Helically Twisted Wire Molecules with Large Effective Conjugation Length

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 π -Conjugated aromatic molecules with helically twisted geometry are fascinating materials with potential applications in inductors, springs, and pressure sensors working on the nanometer scale. In this work, we synthesized π -extended [5], [7], and [9]helicene derivatives 1–3 (Figure 1a).

 π -extended The homogeneously helicenes showed the lowest-energy absorption bands at 500–1100 nm (i.e., $\lambda_{max,abs} = 529, 675$, 809 nm for 1, 2, and 3, respectively), which were attributed to the HOMO→LUMO transitions by DFT calculations. Given that the excitation energy of the original [n] helicenes is largely independent of the length of the helical structure, the prominent red shift of 1-3 in the absorption bands is likely attributed to significant orbital interactions between phenanthrene subunits in the π -extended helicene framework. Enantiomers of 1-3, isolated by HPLC with a chiral column, exhibited Cotton effects with dissymmetry factors of $g_{CD} = 0.2-3.0 \times 10^{-2}$ (Figure 1b). Interestingly, π -extended helicenes 1-3 were not emissive due to their ultrafast nonradiative decay dynamics in the $S_1 \rightarrow S_0$ transition with time constants of about 1 ps, which is marked contrast to emissive planar molecules such as rylenes^[1] and carbon-bridged oligo(phenylene vinylene) molecules^[2].



Figure 1. (a) Molecular structures of π -extended helicenes 1–3 (b) Absorption and circular dichroism (CD) spectra of 1 (top), 2 (middle), 3 (bottom). Red and blue lines denote the (*P*)- and (*M*)-isomers.

[1] A. Bohnen *et al.*, *Angew. Chem. Int. Ed. Engl.*, **1990**, *29*, 525.

[2] M. Morales-Vidal et al., Nat. Commun. 2015, 6, 8458.