## Influence on Counter Anion of Cyclic Thioamidium Salts by Introduction Oxygen Functional Substituent in Iodine-mediated Cyclization of o-Ethynylthiobenzamide

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We have reported the formation of cyclic thioamidium salts using iodine-mediated cyclization reaction with o-ethynylthiobenzamides (Scheme 1) [1]. The cyclic compounds possess iodide (I) as a counter anion. But we also revealed the variation of

counter anion ( $I^{-}$  and/or  $I_{3}^{-}$ ) in the case of thiazole rings instead of thioamide moiety of 1 [2]. Herein we will report the reaction of the substrates with oxygen functional group on benzene ring of 1.

When **1b** bearing methoxy group was treated with 1 mol-equivalent of  $I_2$ , the cyclic thioamidium  $(2b^+)$  was obtained as the precipitated in 65% yield with ratio of  $I^-$ :  $I_3^-$ = 32 : 68 as anion parts (Table 1, Entry 2). The yield and the ratio of anion were increased by using 2

mol-equivalent of I<sub>2</sub> (Entry 3). However, the ratio of counter anion decreased in the case of the substrate bearing acetoxy group (1c) (Entry 4). Thus, the ration of counter anion would be affected by the electronic character of the substituent. And the electron-donating

Table 1. Reaction of 1 with Iodine



Scheme 1

NMe<sub>2</sub>

Ph

l<sub>2</sub> (1 mol-equiv.)

CHCl<sub>3</sub>

r.t., 24 h

NMe<sub>2</sub>

a) Yield and ration were determined by the estimation with weight of the precipitate and the elemental analysis.



Figure 1. Single-crystal X-ray structures of a)  $2b^+ \cdot I_3^-$  and b)  $2c^+ \cdot I^-$ 

substituents enhance the formation of the product with  $I_3$  counter anion. It would be caused by the stabilization ability of the substituent against the cyclic cation part. The structures of  $2b^+ \cdot I_3$  and  $2c^+ \cdot I^-$  were clarified by single-crystal X-ray analysis (Figure 1).

[1] S. Matsumoto, D. Takada, H. Kageyama, M. Akazome, Tetrahedron Lett., 55, 1082 (2014).

[2] S. Matsumoto, R. Sumida, S. E. Tan, M. Akazome, Heterocycles, 93, 755 (2018).