Redox Conduction in 2D Bis(terpyridine)metal(II) Complex Polymers

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Electron transport in thin films is a fundamental phenomenon and plays an important role in electronic/electrochemical devices. Coordination polymers woven up with metal-terpyridine coordination bonds are such polymers featuring robust redox reactions, expected for electrochemical catalysts and electrochromic materials. Previously, we demonstrated electrochromic 2D bis(terpyridine)iron(II) and cobalt(II) coordination polymers, **M-tpy** (M = Fe & Co), with a three-armed terpyridine ligand.¹ For further understanding of electrochemical behaviors of the 2D polymer, here we report the results of potential-dependent conductivity measurements of **Fe-tpy** and **Co-tpy**, and give an explanation for electron transport through bis(terpyridine)iron(II) and cobalt(II) heterolayers (Fig.1).

Potential-dependent conductivity measurements were performed with interdigitated array microelectrodes (IDAs) in an Ar-filled glove box. The results indicated that the conductivity was at local maxima at the mixed-valence states. The conductivity-potential curves were well consistent with the expectations from the electron self-exchange theory, which confirms that redox conduction by electron hopping between $[M(tpy)_2]$ units is dominant in **M-tpy**. By utilizing this electron transport mechanism, we demonstrate unique dissymmetric electron transport phenomena through **Fe-tpy/Co-tpy** vertical junctions which we have recently reported.²



Fig. 1. Redox conductivity in M-tpy measured with IDAs.

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